



$$0.9674 \text{ g } Na_2CO_3 \times \frac{1 \text{ mol } Na_2CO_3}{105.9884 \pm 0.0007 \text{ g}} = 9.127 \times 10^{-3} \text{ mol} \pm 0.093\%$$

$$9.127 \times 10^{-3} \text{ mol } Na_2CO_3 \times \frac{2 \text{ mol } HCl}{1 \text{ mol } Na_2CO_3} = 0.0182 \text{ mol } HCl \pm 0.093\%$$

$$0.0182 \text{ mol } HCl \times \frac{1}{0.02735 \text{ L}} = 0.6654 \pm 0.0015 \frac{\text{mol } HCl}{\text{L}}$$

$$0.667 \pm 0.001 \text{ M } HCl$$

e	% e	value	% e ²
0.0009	0.093		8.649×10^{-3}
0.0007	6.604×10^{-4}		4.362×10^{-7}
	0.093		8.649×10^{-3}
0.0004			
4×10^{-5}	0.146		0.021
0.0015	0.172		0.0296