$2.500\,g \times \frac{2.5\%}{100\%} = 0.0625\,g$

2.500g
0.0625g

$0.0625\,g\ Cu \times \frac{1000\,mg}{g} \times \frac{\quad}{0.1\,L} = 625\,\frac{mg}{L}$ or 625 ppm Cu

625 ppm Cu

$M_1V_1 = M_2V_2$

$625\ ppm \cdot X = 2.5\ ppm\ Cu \cdot 0.1\,L$

$X = 0.4\ ml$

0.4ml

$M_1V_1 = M_2V_2$

$100\ ppm\ Cu \quad X = 1\ ppm\ Cu \cdot 100\,mL$

$X = 1\ mL$

$100\ ppm\ Cu \quad X = 2\ ppm\ Cu \cdot 100\,mL$

$X = 2\,mL$

1mL
2mL
3mL
4mL
5mL

$\frac{50\,mg}{0.1\,L} = 500\,\frac{mg}{L}$ or 500 ppm Cu

500 ppm Cu

$M_1V_1 = M_2V_2$

$500\ ppm \quad X = 5\ ppm \cdot 100\,mL$

$X = 1\,mL$